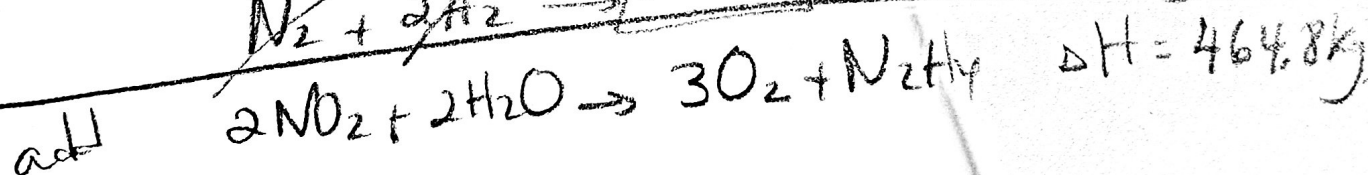
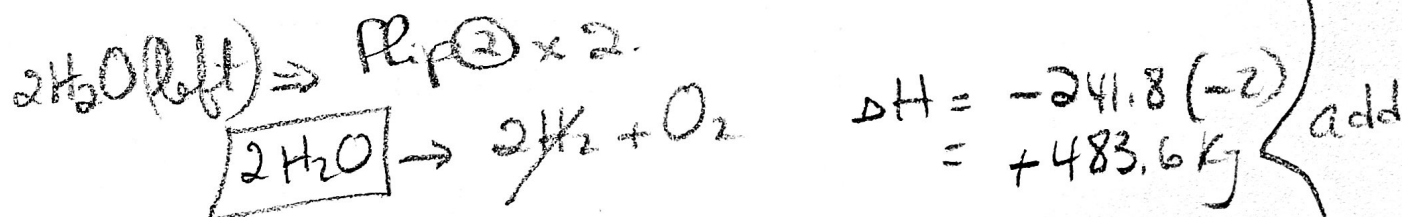
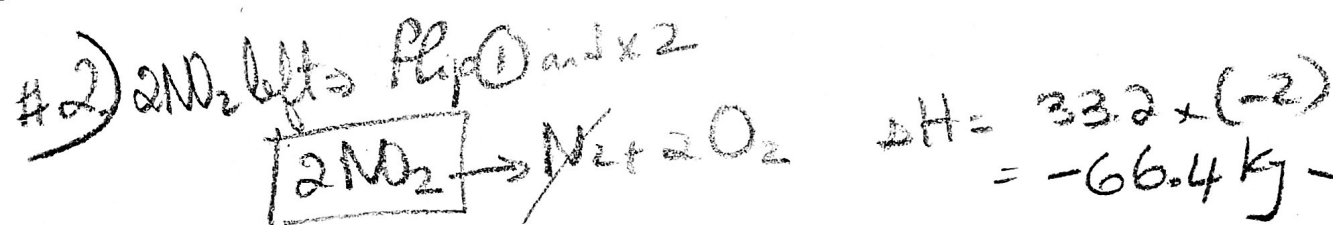
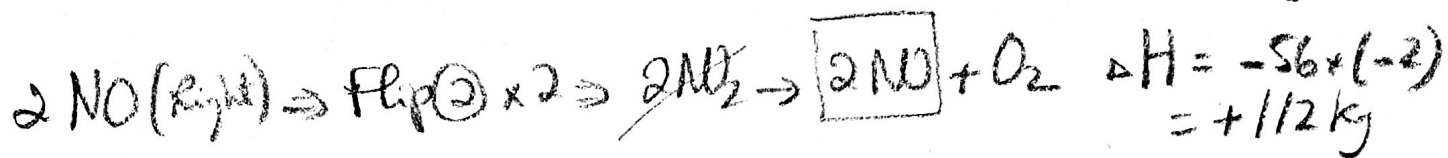
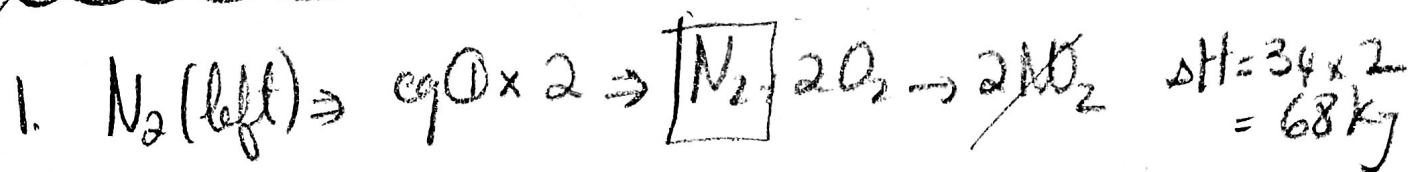
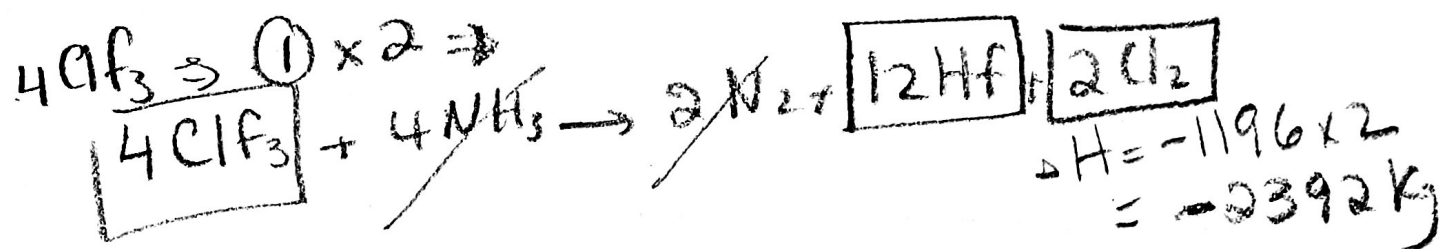
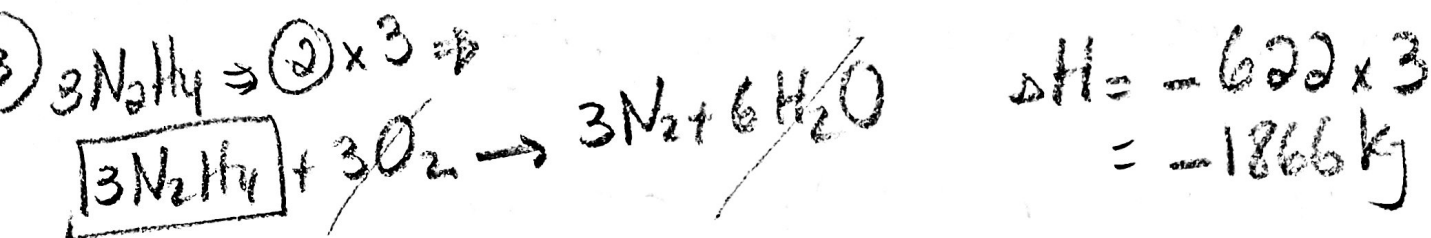


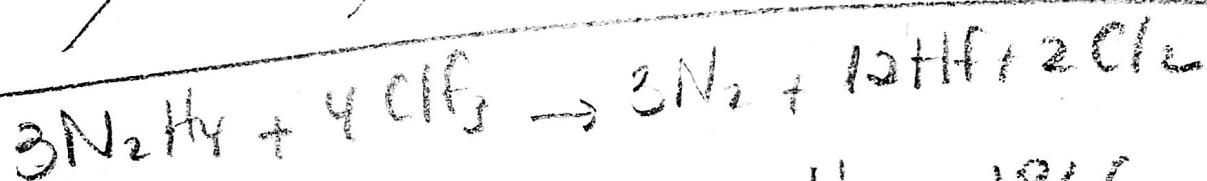
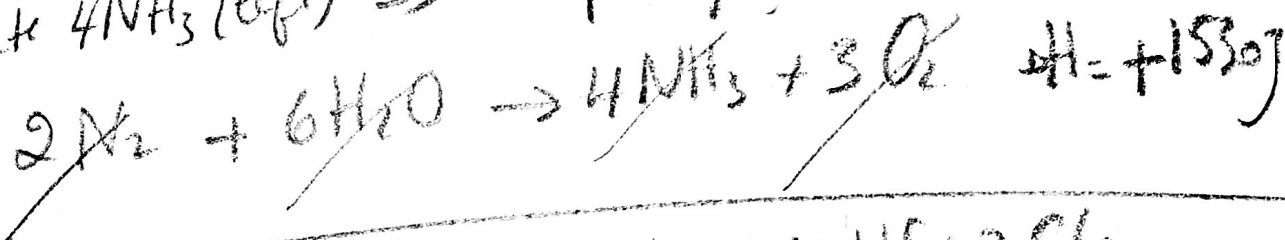
# Practice Problems Hess's Law Notes





12HF: {already there.  
2Cl<sub>2</sub>}

Eliminate 4NH<sub>3</sub> (b.f.)  $\Rightarrow$  flip equation (c)



$$\Delta H = -1866$$

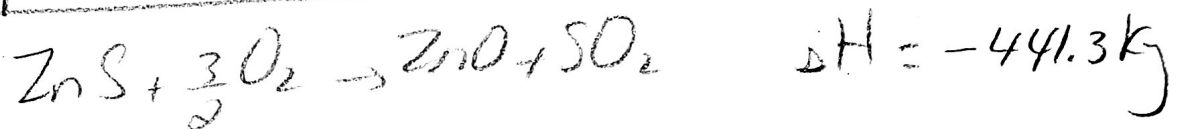
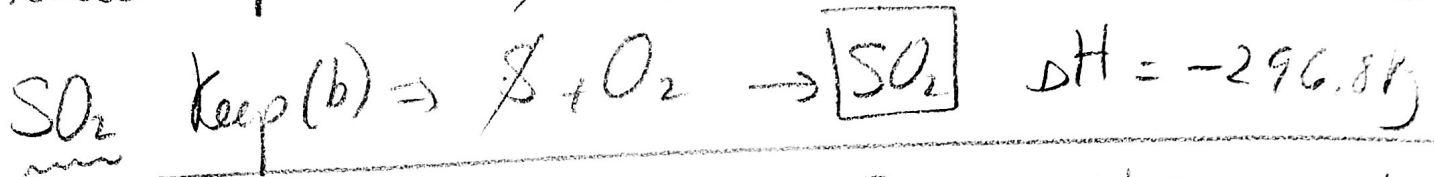
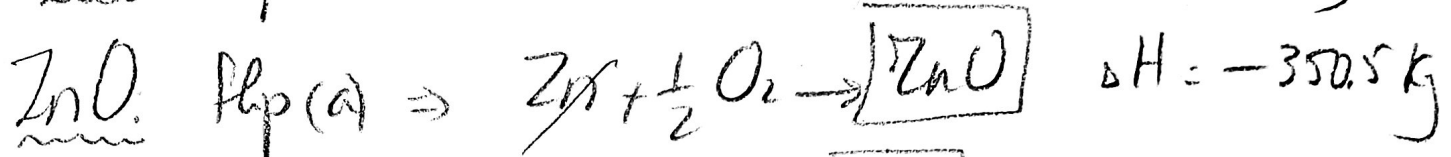
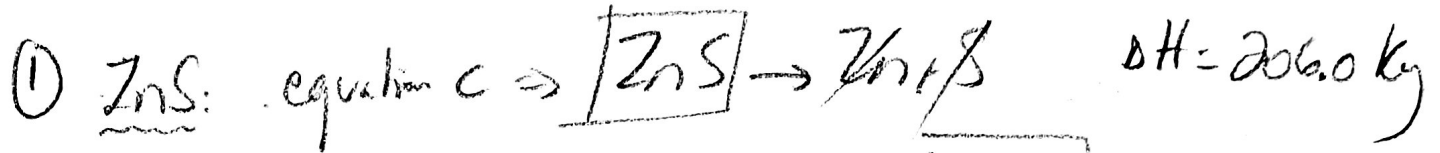
$$-2392$$

$$+1530$$

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$$= -2728 \text{ kJ}$$

- #4) ① Find  $\Delta H$  reaction.  
 ② then  $\Delta H$  for 50.0g ZnS.



$\left\{ \frac{-441.3 \text{ kJ}}{1 \text{ mol ZnS}} = \frac{x \text{ kJ}}{0.513 \text{ mol}} \right\} \Rightarrow x = -441.3 \times 0.513$   
 $= -226 \text{ kJ}$

$50.0 \text{ g ZnS} \times \frac{1 \text{ mol}}{97.46 \text{ g}} = 0.513 \text{ mol}$

$\begin{array}{r} 65.39 \\ + 32.07 \\ \hline 97.46 \end{array}$